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10.1: Energy Transfer (Practice on pg. 342)
1. 97 J 2. 1.3 mol 3. 2.2 10² J 4. 197 K (Section Review on pg. 344)
9. 29.2 kJ 10. 14.6 kJ 11. 0.52 mol 12. 358 K 13. 295 K 14. 0.234 J (K g) 15. 0.864 J (K g) 18. 100 J (K mol) 19. 0.6 J (K g)
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Chapter 14: Chemical Kinetics (pg. 489 to 493)
27. Activation energy is the amount of energy needed to overcome for the products to form. All reactions have some kind of activation energy. The higher is the activation energy, the slower the reaction rate.
30. Methane and oxygen gas mixture does not react spon...

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1. System: - any specific part of the universe that is of interest to us. Surrounding: - the rest of the universe outside the system. Open System: - a system that can exchange mass and energy (usually in the form of heat) with its sur...





















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