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TABLE 4.1 REVISED 1997 BY C. G. WOHL, LBNL. THE ATOMIC NUMBER TOP LEFT IS THE NUMBER OF PROTONS IN THE NUCLEUS. THE ATOMIC MASS BOTTOM IS WEIGHTED BY

isotopic abundances in the Earth's surface. Atomic masses are relative to the mass of the carbon-12 isotope defined to be exactly 12 unified atomic mass units. Errors range from 1 to 9 in the last digit quoted. Relative isotopic abundances often vary considerably both in natural and commercial samples. A number in parentheses is the mass of the longest-lived isotope of that element. No stable isotope exists...

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ATOMIC STRUCTURE NAME ~~~~~

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many ele...

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..... ~-ATOMIC MASS. Elements come in a variety of isotopes, meaning they are made up of atoms with the same atomic number but different atomic masses. These atoms differ in the number of neutrons. The average atomic mass is the weighted average of all the isotopes of an element. Example: A sample of cesium is 75% ¹³³Cs, 20% ¹³²Cs, and 5% ¹³⁴Cs. What is ...

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Chemistry: Form LS2.4A Name ATOMS Date Period Average Atomic Mass Aim to calculate the average mass of an element. Notes Atomic mass. The relative mass of an atom is the sum of its protons and neutrons. q Examples p Carbon 12 has 6 protons and 6 neutrons (6 + 6 = 12) p Carbon 14 has 6 protons and 8 neutrons (6 + 8 = 14) q The relative mass of any isotope is an integer. Average atomic mass. The masses of ...

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